

Text Book Of Physical Chemistry

This is a new undergraduate textbook on physical chemistry by Horia Metiu published as four separate paperback volumes. These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research. By u

The advancements in society are intertwined with the advancements in science. To understand how changes in society occurred, and will continue to change, one has to have a basic understanding of the laws of physics and chemistry. Physical Chemistry: Multidisciplinary Applications in Society examines how the laws of physics and chemistry (physical chemistry) explain the dynamic nature of the Universe and events on Earth, and how these events affect the evolution of society (multidisciplinary applications). The ordering of the chapters reflects the natural flow of events in an evolving Universe: Philosophy of Science, the basis of the view that natural events have natural causes - Cosmology, the origin of everything from the Big Bang to the current state of the Universe - Geoscience, the physics and chemistry behind the evolution of the planet Earth from its birth to the present - Life Science, the molecules and mechanisms of life on Earth - Ecology, the interdependence of all components within the Ecosphere and the Universe - Information Content, emphasis on how words and phrases and framing of issues affect opinions, reliability of sources, and the limitations of

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knowledge. Addresses the four Ws of science: Why scientists believe Nature works the way it does, Who helped develop the fields of science, What theories of natural processes tell us about the nature of Nature, and Where our scientific knowledge is taking us into the future Gives a historical review of the evolution of science, and the accompanying changes in the philosophy of how science views the nature of the Universe Explores the physics and chemistry of Nature with minimal reliance on mathematics Examines the structure and dynamics of the Universe and our Home Planet Earth Provides a detailed analysis of how humans, as members of the Ecosphere, have influenced, and are continuing to influence, the dynamics of events on the paludarium called Earth Presents underlying science of current political issues that shape the future of humankind Emphasizes how words and phrases and framing of issues can influence the opinions of members of society Makes extensive use of metaphors and everyday experiences to illustrate principles in science and social interactions

With its easy-to-read approach and focus on core topics, PHYSICAL CHEMISTRY, 2e provides a concise, yet thorough examination of calculus-based physical chemistry. The Second Edition, designed as a learning tool for students who want to learn physical chemistry in a functional and relevant way, follows a traditional organization and now features an increased focus on thermochemistry, as well as new problems, new two-column examples, and a dynamic new four-color design. Written by a dedicated chemical educator and

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researcher, the text also includes a review of calculus applications as applied to physical chemistry. Important Notice: Media content referenced within the product description or the product text may not be available in the ebook version.

aspects of the learning process are fully supported, including the understanding of terminology, notation, mathematical concepts, and the application of physical chemistry to other branches of science." "Building on the heritage of the world-renowned Atkins' Physical Chemistry , Quanta, Matter, and Change gives a refreshing new insight into the familiar by illuminating physical chemistry from a new direction." --Book Jacket.

This book is the fifth of the six-volume series, which provides an extensive coverage of Physical Chemistry.

Each volume includes a large number of illustrative numericals and typical problems to highlight the principles involved. IUPAC recommendations and SI units have been adopted throughout. The present book describes Adsorption, Chemical Kinetics,

Photochemistry, Statistical hermodynamics, and Macromolecules. A new chapter on Introduction to Irreversible Processes has been added. Salient

Features: • Comprehensive coverage to adsorption, chemical kinetics, photochemistry, statistical thermodynamics, macromolecules • Emphasis given to applications and principles • Explanation of equations in the form of solved problems and numericals • IUPAC recommendations and SI units have been adopted throughout. • Rich and illustrious pedagogy

Written primarily to meet the requirements of students at

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the undergraduate level, this book aims for a self-learning approach. The fundamentals of physical chemistry have been explained with illustrations, diagrams, tables, experimental techniques and solved problems.

Exploring the structure and physical and chemical properties of solutions, dispersions, soft solids, fats, and cellular systems, *Physical Chemistry of Foods* describes the physiochemical principles of the reactions and conversions that occur during the manufacture, handling, and storage of foods. Coverage progresses from aspects of thermodynamics, bonds and interaction forces, and reaction kinetics, to transport phenomena, polymers, colloidal interactions, nucleation, glass transitions and freezing, and soft solids. This comprehensive volume effectively clarifies the physicochemical processes encountered in food product development.

Textbook of Physical Chemistry, together with the companion books on organic chemistry and inorganic chemistry, meets the complete requirements of undergraduate students of chemistry across India. In a book comprising all the classical topics which span physical chemistry including chemical kinetics, electrochemistry and thermodynamics among others, uniformity in the depth of coverage of each topic is not easy to attain in view of the disjointed pace of growth of each discipline. Nevertheless, care has been taken to ensure that the material in this book will sustain the interest of students and motivate them to learn

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physical chemistry. In order to aid students, every chapter contains the Objectives at the beginning and Key Points at the end. Various aspects of physical chemistry are dealt with in a lucid manner and interesting related matter is highlighted in boxes. The derivations are given in a comprehensible manner. Since physical chemistry involves numericals, several worked examples complement the text. The exercises at the end of each chapter, in particular, will be extremely valuable to sharpen the problem-solving skills and direct the student towards appreciating the nuances of physical chemistry. The first part of this book looks at the consequence of chemical and topological defects existing on real surfaces, which explain the wettability of super hydrophilic and super hydrophobic surfaces. There follows an in-depth analysis of the acido-basicity of surfaces with, as an illustration, different wettability experiments on real materials. The next chapter deals with various techniques enabling the measurement of acido basicity of the surfaces including IR and XPS technics. The last part of the book presents an electrochemical point of view which explains the surface charges of the oxide at contact with water or other electrolyte solutions in the frame of Bronsted acido-basicity concept. Various consequences are deduced from such analyses illustrated by original measurement of the point of zero charge or by understanding the basic

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principles of the electrowetting experiments.

This textbook covers the fundamentals of physical chemistry, explaining the concepts in an accessible way and guiding the readers in a step-by-step manner. The contents are broadly divided into two sections: the classical physico-chemical topics (thermodynamics, kinetics, electrochemistry, transport, and catalysis), and the fabric of matter and its interactions with radiation. Particular care has been taken in the presentation of the algebraic parts of physico-chemical concepts, so that the readers can easily follow the explanations and re-work relevant discussion and derivations with pen and paper. The book is accompanied by a rich mathematical appendix. Each chapter includes a selection of (numerical) exercises and problems, so that students can practice and apply the learned topics. An appendix with solutions allows for controlling the learning success. Carefully prepared illustrative color images make this book a great support for teaching physical chemistry to undergraduate students. This textbook mainly addresses undergraduate students in life sciences, biochemistry or engineering, offering them a comprehensive and comprehensible introduction for their studies of physical chemistry. It will also appeal to undergraduate chemistry students as an accessible introduction for their physical chemistry studies.

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A Textbook of Physical Chemistry, Second Edition serves as an introductory text to physical chemistry. Topics covered range from wave mechanics and chemical bonding to molecular spectroscopy and photochemistry; ideal and nonideal gases; the three laws of thermodynamics; thermochemistry; and solutions of nonelectrolytes. The kinetics of gas-phase reactions; colloids and macromolecules; and nuclear chemistry and radiochemistry are also discussed. This edition is comprised of 22 chapters; the first of which introduces the reader to the behavior of ideal and nonideal gases, with particular emphasis on the van der Waals equation. The discussion then turns to the kinetic molecular theory of gases and the application of the Boltzmann principle to the treatment of molar polarization; dipole and magnetic moments; the phenomenology of light absorption; and classical and statistical thermodynamics. The chapters that follow focus on the traditional sequence of chemical and phase equilibria, electrochemistry, and chemical kinetics in gas phase and solution phase. This book also considers wave mechanics and its applications; molecular spectroscopy and photochemistry; and the excited state, and then concludes with an analysis of crystal structure, colloid and polymer chemistry, and radio and nuclear chemistry. This reference material is intended primarily as an introductory text for students of physical chemistry.

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This book is the last of the seven-volume series, which provides an extensive coverage of several topics of Physical Chemistry. Volume 7 has been designed to assist the students to understand theory, procedures, etc. of an experiment before they actually perform it in the laboratory. Wherever possible, analysis of data of the experiment obtained by a student is added as an annexure to the experiment. This also includes the computational analysis involving leastsquare fitting of data of the experiment.

Coverage of Physical Chemistry. Each volume includes a large number of illustrative numericals and typical problems to highlight the principles involved. IUPAC recommendations and SI units have been adopted throughout. The present book describes Wave Mechanics, Energy Quantization and Atomic Structure, Theories of Covalent Bond, Electrical and Magnetic Properties of Molecules, Molecular Spectroscopy, Molecular Symmetry and its Applications. Salient Features: • Comprehensive coverage of wave mechanics, energy quantization and atomic structure, theories of covalent bond, electrical and magnetic properties of molecules, molecular spectroscopy, molecular symmetry and its applications • Emphasis given to applications and principles • Explanation of equations in the form of solved problems and numericals • IUPAC recommendations and SI units have been adopted

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throughout • Rich and illustrious pedagogy

Atkins' Physical Chemistry: Molecular Thermodynamics and Kinetics is designed for use on the second semester of a quantum-first physical chemistry course. Based on the hugely popular Atkins' Physical Chemistry, this volume approaches molecular thermodynamics with the assumption that students will have studied quantum mechanics in their first semester. The exceptional quality of previous editions has been built upon to make this new edition of Atkins' Physical Chemistry even more closely suited to the needs of both lecturers and students. Re-organised into discrete 'topics', the text is more flexible to teach from and more readable for students. Now in its eleventh edition, the text has been enhanced with additional learning features and maths support to demonstrate the absolute centrality of mathematics to physical chemistry. Increasing the digestibility of the text in this new approach, the reader is brought to a question, then the math is used to show how it can be answered and progress made. The expanded and redistributed maths support also includes new 'Chemist's toolkits' which provide students with succinct reminders of mathematical concepts and techniques right where they need them. Checklists of key concepts at the end of each topic add to the extensive learning support provided throughout the book, to reinforce the main take-home messages in each section. The

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coupling of the broad coverage of the subject with a structure and use of pedagogy that is even more innovative will ensure Atkins' Physical Chemistry remains the textbook of choice for studying physical chemistry.

This book is ideal for use in a one-semester introductory course in physical chemistry for students of life sciences. The author's aim is to emphasize the understanding of physical concepts rather than focus on precise mathematical development or on actual experimental details. Subsequently, only basic skills of differential and integral calculus are required for understanding the equations. The end-of-chapter problems have both physiochemical and biological applications.

This comprehensive textbook, now in its second edition, is mainly written as per the latest syllabi of physical chemistry of all the leading universities of India as well as the new syllabus recommended by the UGC. This thoroughly revised and updated edition covers the principal areas of physical chemistry, such as thermodynamics, quantum chemistry, molecular spectroscopy, chemical kinetics, electrochemistry and nanotechnology. In a methodical and accessible style, the book discusses classical, irreversible and statistical thermodynamics and statistical mechanics, and describes macroscopic chemical systems, steady states and thermodynamics at a molecular level. It elaborates the underlying principles of quantum mechanics, molecular spectroscopy, X-ray crystallography and solid state chemistry along with their applications. The book explains various instrumentation techniques such as

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potentiometry, polarography, voltammetry, conductometry and coulometry. It also describes kinetics, rate laws and chemical processes at the electrodes. In addition, the text deals with chemistry of corrosion and nanomaterials. This text is primarily designed for the undergraduate and postgraduate students of chemistry (B.Sc. and M.Sc.) for their course in physical chemistry. Key Features

- Gives a thorough treatment to ensure a solid grasp of the material.
- Presents a large number of figures and diagrams that help amplify key concepts.
- Contains several worked-out examples for better understanding of the subject matter.
- Provides numerous chapter-end exercises to foster conceptual understanding.

This elegant book provides a student-friendly introduction to the subject of physical chemistry. It is concise and more compact than standard textbooks on the subject and it emphasises the two important concepts underpinning physical chemistry: quantum mechanics and the second law of thermodynamics. The principles are challenging to students because they both focus on uncertainty and probability. The book explains these fundamental concepts clearly and shows how they offer the key to understanding the wide range of chemical phenomena including atomic and molecular spectra, the structure and properties of solids, liquids and gases, chemical equilibrium, and the rates of chemical reactions.

Designed for a two-semester introductory course sequence in physical chemistry, *Physical Chemistry: A Modern Introduction, Second Edition* offers a streamlined

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introduction to the subject. Focusing on core concepts, the text stresses fundamental issues and includes basic examples rather than the myriad of applications often presented in other, more encyclopedic books. Physical chemistry need not appear as a large assortment of different, disconnected, and sometimes intimidating topics. Instead, students should see that physical chemistry provides a coherent framework for chemical knowledge, from the molecular to the macroscopic level. The book offers: Novel organization to foster student understanding, giving students the strongest sophistication in the least amount of time and preparing them to tackle more challenging topics Strong problem-solving emphasis, with numerous end-of-chapter practice exercises, over two dozen in-text worked examples, and a number of clearly identified spreadsheet exercises A quick review in calculus, via an appendix providing the necessary mathematical background for the study of physical chemistry Powerful streamlined development of group theory and advanced topics in quantum mechanics, via appendices covering molecular symmetry and special quantum mechanical approaches This textbook presents a straightforward introduction to physical chemistry. Whilst stressing the fundamentals of the subject, it avoids the mathematical details of specialised techniques such as quantum theory, nuclear magnetic resonance, and spectroscopy. In order to promote an appreciation of 3-dimensional structure in the study of stereo-chemistry and solids, many of the illustrations are presented as stereoscopic views, and directions for observing them are given in an appendix.

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Each chapter ends with a set of problems of varying degrees of difficulty, which will assist the student in gaining familiarity with the themes of the book, and in testing their ability to apply these themes to new situations; full solutions are provided. The SI system of units is used throughout and appendices serve as a useful reference source of numerical data. Some mathematical arguments are also developed in appendices, because their inclusion in the text might distract readers from the development of the subject. The book has been developed from an earlier publication by the authors entitled *Modern Physical Chemistry*, published by Penguin Books Ltd.

This book has been successfully guiding undergraduate students of science, engineering and pharmacy of the Indian universities since 1978 due to its approach of teaching the subject in the simplest possible way. The book emphasizes on fundamental rather than excessive details and develops the topics from the first principles. It contains a considerable number of worked-out examples exposing the students to practical applications of equations and helping them comprehend the magnitude of many different physicochemical quantities. Both the traditional cgs/esu and the newer SI systems of units have been used identically. This is so because in spite of wider acceptance of the SI units, the cgs units continue to be used in most chemical literature. New in this Edition • Quick Recap' section with every chapter to bring the concepts on fingertips • Vastly augmented section on MCQs for complete comprehension • Additional review questions to make them broad based • Revised and

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updated topics

Defects play an important role in determining the properties of solids. This book provides an introduction to chemical bond, phonons, and thermodynamics; treatment of point defect formation and reaction, equilibria, mechanisms, and kinetics; kinetics chapters on solid state processes; and electrochemical techniques and applications. * Offers a coherent description of fundamental defect chemistry and the most common applications. * Up-to-date trends and developments within this field. * Combines electrochemical concepts with aspects of semiconductor physics.

A textbook for B.Sc Classes as per the UGC Model Syllabus. The book is visually beautiful and authors communicate their enthusiasm and enjoyment of the subject in every chapter. This textbook is currently in use at hundreds of colleges and universities throughout the country and is a national best-seller. There are hundreds of computer-generated coloured diagrams, graphs, photos and tables .

For many processes and applications in science and technology a basic knowledge of liquids and solutions is a must. Gaining a better understanding of the behavior and properties of pure liquids and solutions will help to improve many processes and to advance research in many different areas. This book provides a comprehensive, self-contained and integrated survey of this topic and is a must-have for many chemists, chemical engineers and material scientists, ranging from newcomers in the field to more experienced researchers. The author offers a clear, well-structured didactic

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approach and provides an overview of the most important types of liquids and solutions. Special topics include chemical reactions, surfaces and phase transitions. Suitable both for introductory as well as intermediate level as more advanced parts are clearly marked. Includes also problems and solutions.

This text on Physical Chemistry encourages users to understand rather than just rote learning. Simple and lucid language has been used to explain principles and practices. The coverage spans large variety of topics including Gaseous State, Liquid State, Solid State, Colloidal State, Chemical Kinetics, Thermodynamics, Chemical Equilibrium, Distribution Law, Electrochemistry, Atomic Structure and Wave Mechanics, Spectroscopy, Photochemistry, Polymers and Nuclear Chemistry.

Salient Features : • Latest IUPAC notations and SI units used • More than 1200 stepwise solved problems and numerical problems • Special features such as interesting facts, common pitfalls, key terms

A Textbook of Physical Chemistry: Second Edition provides both a traditional and theoretical approach in the study of physical chemistry. The book covers subjects usually covered in chemistry textbooks such as ideal and non-ideal gases, the kinetic molecular theory of gases and the distribution laws, and the additive physical properties of matter. Also covered are the three laws of thermodynamics, thermochemistry, chemical equilibrium, liquids and their simple phase equilibria, the solutions of nonelectrolytes, and heterogenous equilibrium. The text is recommended for college-level chemistry students, especially those who are in need of a textbook for the

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subject.

This book is designed for a one-semester course, for undergraduates, not necessarily chemistry majors, who need to know something about physical chemistry. The emphasis is not on mathematical rigor, but subtleties and conceptual difficulties are not hidden. It covers the essential topics in physical chemistry, including the state of matter, thermodynamics, chemical kinetics, phase and chemical equilibria, introduction to quantum theory, and molecular spectroscopy. Supplementary materials are available upon request for all instructors who adopt this book as a course text. Please send your request to sales@wspc.com.

Emphasizes a molecular approach to physical chemistry, discussing principles of quantum mechanics first and then using those ideas in development of thermodynamics and kinetics. Chapters on quantum subjects are interspersed with ten math chapters reviewing mathematical topics used in subsequent chapters. Includes material on current physical chemical research, with chapters on computational quantum chemistry, group theory, NMR spectroscopy, and lasers. Units and symbols used in the text follow IUPAC recommendations. Includes exercises. Annotation copyrighted by Book News, Inc., Portland, OR

A leading book for 80 years, Silbey's Physical Chemistry features exceptionally clear explanations of the concepts and methods of physical chemistry for students who have had a year of calculus and a year of physics. The basic theory of chemistry is presented from the viewpoint of academic physical chemists, but the many practical

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applications of physical chemistry are integrated throughout the text. The problems in the text also reflect a skillful blend of theory and practical applications. This text is ideally suited for a standard undergraduate physical chemistry course taken by chemistry, chemical engineering, and biochemistry majors in their junior or senior year.

An advanced-level textbook of physical chemistry for the graduate (B.Sc) and postgraduate (M.Sc) students of Indian and foreign universities. This book is a part of four volume series, entitled "A Textbook of Physical Chemistry – Volume I, II, III, IV". CONTENTS: Chapter 1. Quantum Mechanics – I: Postulates of quantum mechanics; Derivation of Schrodinger wave equation; Max-Born interpretation of wave functions; The Heisenberg's uncertainty principle; Quantum mechanical operators and their commutation relations; Hermitian operators (elementary ideas, quantum mechanical operator for linear momentum, angular momentum and energy as Hermitian operator); The average value of the square of Hermitian operators; Commuting operators and uncertainty principle(x & p ; E & t); Schrodinger wave equation for a particle in one dimensional box; Evaluation of average position, average momentum and determination of uncertainty in position and momentum and hence Heisenberg's uncertainty principle; Pictorial representation of the wave equation of a particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level; Lowest energy of the particle. Chapter 2. Thermodynamics – I: Brief resume of first and second Law of thermodynamics; Entropy changes in reversible and irreversible processes; Variation of entropy with temperature, pressure and volume; Entropy concept as a measure of unavailable energy and criteria for the

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spontaneity of reaction; Free energy, enthalpy functions and their significance, criteria for spontaneity of a process; Partial molar quantities (free energy, volume, heat concept); Gibb's-Duhem equation. Chapter 3. Chemical Dynamics – I: Effect of temperature on reaction rates; Rate law for opposing reactions of 1st order and 2nd order; Rate law for consecutive & parallel reactions of 1st order reactions; Collision theory of reaction rates and its limitations; Steric factor; Activated complex theory; Ionic reactions: single and double sphere models; Influence of solvent and ionic strength; The comparison of collision and activated complex theory. Chapter 4. Electrochemistry – I: Ion-Ion Interactions: The Debye-Huckel theory of ion-ion interactions; Potential and excess charge density as a function of distance from the central ion; Debye Huckel reciprocal length; Ionic cloud and its contribution to the total potential; Debye - Huckel limiting law of activity coefficients and its limitations; Ion-size effect on potential; Ion-size parameter and the theoretical mean-activity coefficient in the case of ionic clouds with finite-sized ions; Debye - Huckel-Onsager treatment for aqueous solutions and its limitations; Debye-Huckel-Onsager theory for non-aqueous solutions; The solvent effect on the mobility at infinite dilution; Equivalent conductivity (Λ) vs. concentration $c^{1/2}$ as a function of the solvent; Effect of ion association upon conductivity (Debye- Huckel - Bjerrum equation). Chapter 5. Quantum Mechanics – II: Schrodinger wave equation for a particle in a three dimensional box; The concept of degeneracy among energy levels for a particle in three dimensional box; Schrodinger wave equation for a linear harmonic oscillator & its solution by polynomial method; Zero point energy of a particle possessing harmonic motion and its consequence; Schrodinger wave equation for three dimensional Rigid rotator; Energy of rigid rotator; Space quantization; Schrodinger wave equation for hydrogen atom,

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separation of variable in polar spherical coordinates and its solution; Principle, azimuthal and magnetic quantum numbers and the magnitude of their values; Probability distribution function; Radial distribution function; Shape of atomic orbitals (s, p & d). Chapter 6. Thermodynamics – II: Clausius-Clayperon equation; Law of mass action and its thermodynamic derivation; Third law of thermodynamics (Nernst heat theorem, determination of absolute entropy, unattainability of absolute zero) and its limitation; Phase diagram for two completely miscible components systems; Eutectic systems, Calculation of eutectic point; Systems forming solid compounds $A_x B_y$ with congruent and incongruent melting points; Phase diagram and thermodynamic treatment of solid solutions. Chapter 7. Chemical Dynamics – II: Chain reactions: hydrogen-bromine reaction, pyrolysis of acetaldehyde, decomposition of ethane; Photochemical reactions (hydrogen - bromine & hydrogen -chlorine reactions); General treatment of chain reactions (ortho-para hydrogen conversion and hydrogen - bromine reactions); Apparent activation energy of chain reactions, Chain length; Rice-Herzfeld mechanism of organic molecules decomposition (acetaldehyde); Branching chain reactions and explosions (H_2-O_2 reaction); Kinetics of (one intermediate) enzymatic reaction : Michaelis-Menton treatment; Evaluation of Michaelis 's constant for enzyme-substrate binding by Lineweaver-Burk plot and Eadie-Hofstae methods; Competitive and non-competitive inhibition. Chapter 8. Electrochemistry – II: Ion Transport in Solutions: Ionic movement under the influence of an electric field; Mobility of ions; Ionic drift velocity and its relation with current density; Einstein relation between the absolute mobility and diffusion coefficient; The Stokes- Einstein relation; The Nernst -Einstein equation; Walden's rule; The Rate-process approach to ionic migration; The Rate process equation for

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equivalent conductivity; Total driving force for ionic transport, Nernst - Planck Flux equation; Ionic drift and diffusion potential; the Onsager phenomenological equations; The basic equation for the diffusion; Planck-Henderson equation for the diffusion potential.

Thanks to the progress made in instruments and techniques, the methods in physical chemistry have developed rapidly over the past few decades, making them increasingly valuable for scientists of many disciplines. These two must-have volumes meet the needs of the scientific community for a thorough overview of all the important methods currently used. As such, this work bridges the gap between standard textbooks and review articles, covering a large number of methods, as well as the motivation behind their use. A uniform approach is adopted throughout both volumes, while the critical comparison of the advantages and disadvantages of each method makes this a valuable reference for physical chemists and other scientists working with these techniques. Peter Atkins' Very Short Introduction explores the contributions physical chemistry has made to all branches of chemistry. Providing insight into its central concepts Atkins reveals the cultural contributions physical chemistry has made to our understanding of the natural world.

2019 RITA® Award Winner for Contemporary Romance: Mid-Length! After four lousy boyfriends in a row, chemical engineer Penny Popplestone swears off men until she can figure out why they keep cheating on her. But her no-men resolution hits a snag when the mysterious and superhumanly hot barista at her favorite coffee shop strikes up a friendship with her. Penny strives to keep things platonic, but when Caleb gives her the kiss of her life, she realizes he wants to be more than just friends. Tired of always being “good little Penny,” she throws caution to the wind and pursues a no-strings fling with the hottie barista. It’s not like they have

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anything in common beyond scorching physical chemistry, so what does she have to lose? Only her heart. Now, this fanfic-reading, plus-size heroine faces an unsolvable problem. What do you do when being apart is unbearable...but being together is impossible? This steamy, lighthearted romance is the third in a series of standalone rom-coms featuring geeky heroines who work in STEM fields.

This is a new undergraduate textbook on physical chemistry by Horia Metiu published as four separate paperback volumes. These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research. By using the computer to solve problems that include actual experimental data, the author is able to cover the subject matter at a practical level. The books closely integrate the theoretical chemistry being taught with industrial and laboratory practice. This approach enables the student to compare theoretical projections with experimental results, thereby providing a realistic grounding for future practicing chemists and engineers. Each volume of Physical Chemistry includes Mathematica[®] and Mathcad[®] Workbooks on CD-ROM. Metiu's four separate volumes-Thermodynamics, Statistical Mechanics, Kinetics, and Quantum Mechanics-offer built-in flexibility by allowing the subject to be covered in any order. These textbooks can be used to teach physical chemistry without a computer, but the experience is enriched substantially for those students who do learn how to read and write Mathematica[®] or Mathcad[®] programs. A TI-89 scientific calculator can be used to solve most of the exercises and problems.

Much of chemistry is motivated by asking 'How'? How do I make a primary alcohol? React a Grignard reagent with formaldehyde. Physical chemistry is motivated by asking

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'Why'? The Grignard reagent and formaldehyde follow a molecular dance known as a reaction mechanism in which stronger bonds are made at the expense of weaker bonds. If you are interested in asking 'why' and not just 'how', then you need to understand physical chemistry. Physical Chemistry: How Chemistry Works takes a fresh approach to teaching in physical chemistry. This modern textbook is designed to excite and engage undergraduate chemistry students and prepare them for how they will employ physical chemistry in real life. The student-friendly approach and practical, contemporary examples facilitate an understanding of the physical chemical aspects of any system, allowing students of inorganic chemistry, organic chemistry, analytical chemistry and biochemistry to be fluent in the essentials of physical chemistry in order to understand synthesis, intermolecular interactions and materials properties. For students who are deeply interested in the subject of physical chemistry, the textbook facilitates further study by connecting them to the frontiers of research. Provides students with the physical and mathematical machinery to understand the physical chemical aspects of any system. Integrates regular examples drawn from the literature, from contemporary issues and research, to engage students with relevant and illustrative details. Important topics are introduced and returned to in later chapters: key concepts are reinforced and discussed in more depth as students acquire more tools. Chapters begin with a preview of important concepts and conclude with a summary of important equations. Each chapter includes worked examples and exercises: discussion questions, simple equation manipulation questions, and problem-solving exercises. Accompanied by supplementary online material: worked examples for students and a solutions manual for instructors. Written by an experienced instructor, researcher and author in physical chemistry, with a voice and perspective

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that is pedagogical and engaging.

This is the physical chemistry textbook for students with an affinity for computers! It offers basic and advanced knowledge for students in the second year of chemistry masters studies and beyond. In seven chapters, the book presents thermodynamics, chemical kinetics, quantum mechanics and molecular structure (including an introduction to quantum chemical calculations), molecular symmetry and crystals. The application of physical-chemical knowledge and problem solving is demonstrated in a chapter on water, treating both the water molecule as well as water in condensed phases. Instead of a traditional textbook top-down approach, this book presents the subjects on the basis of examples, exploring and running computer programs (Mathematica®), discussing the results of molecular orbital calculations (performed using Gaussian) on small molecules and turning to suitable reference works to obtain thermodynamic data. Selected Mathematica® codes are explained at the end of each chapter and cross-referenced with the text, enabling students to plot functions, solve equations, fit data, normalize probability functions, manipulate matrices and test physical models. In addition, the book presents clear and step-by-step explanations and provides detailed and complete answers to all exercises. In this way, it creates an active learning environment that can prepare students for pursuing their own research projects further down the road. Students who are not yet familiar with Mathematica® or Gaussian will find a valuable introduction to computer-based problem solving in the molecular sciences. Other computer applications can alternatively be used. For every chapter learning goals are clearly listed in the beginning, so that readers can easily spot the highlights, and a glossary in the end of the chapter offers a quick look-up of important terms.

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